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### SOLUTION-MINERAL EQUILIBRIA PART 1: CARBONATES

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The solubility product ( $K_{sp}$ ) is the equilibrium constant of the dissociation reaction of the mineral-containing salt in water. Hence,  $K_{sp}$  is a constant at a given temperature. The solubility ( $S$ ) of a mineral salt is the amount of the salt that is dissolved per unit volume. This quantity may vary, depending on the conditions.

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The book *Solutions, Minerals, and Equilibria* co-authored in 1965 by Garrels and Charles L. Christ revolutionized aqueous geochemistry. Garrels earned a bachelor's degree in geology from The University of Michigan in 1937. He went on to earn an M.S. degree from Northwestern in 1939, his thesis work was on iron ores of Newfoundland in 1938.

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